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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.  $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$   $3 \text{ O}_2 \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}$   $1 \text{ k J mol}^{-1}$   $2 \text{ k J mol}^{-1}$   $8 \text{ k J mol}^{-1}$   $1 \text{ k J mol}^{-1}$   $5 \text{ k J mol}^{-1}$   $6 \text{ k J mol}^{-1}$  (hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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